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Chapter 9: Chemical Reactions in Aqueous Solution 9.12 Ionic compounds, strong acids, and strong bases (metal hydroxides) are strong electrolytes (completely broken up into ions of the compound). Weak acids and weak bases are weak electrolytes. Molecular substances other than acids or bases are nonelectrolytes. a. nonelectrolyte (Water is actually a very weak electrolyte.)

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What is the molarity of a solution prepared by dissolving 36.0 g of NaOH in enough water to make 1.50 L of solution? 0.600 M How many moles of HCl are present in 35.5 mL of a 0.200 M solution?

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