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FOR HEAT CAPACITY 2  $T = \frac{1}{k} \left( \frac{U}{N} \right)^{1/3}$  (7)  
 $\frac{1}{T} = \frac{1}{k} \left( \frac{U}{N} \right)^{-2/3} \frac{dU}{dN}$  (8)  $= k \left[ \ln(q+N) + 1 \ln q \right]$  (9)  
 $= k \ln q + N \frac{1}{q}$  (10)  $= k \ln U + N \frac{1}{U}$  (11)  $T = \frac{1}{k \ln U + N \frac{1}{U}}$  (12) We can invert this to obtain  $U(T)$ :  $e^{-kT} = \frac{U}{U+N} = \frac{U}{N} \left( 1 + \frac{U}{N} \right)^{-1}$  (13)  $U = N \frac{e^{-kT}}{1 + e^{-kT}}$  (14) Finally, the heat capacity is obtained from  $C = \frac{dU}{dT}$  (15)  $= N e^{-kT}$

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$$=kT e =kT \frac{1}{2} kT^2 (16) = 2Ne =kT \frac{1}{2} kT^2 e$$
$$=kT \frac{1}{2} (17)$$

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Daniel V. Schroeder Department of  
Physics Weber State University.

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The thermal energy of a monatomic gas is  $\frac{3}{2} NkT = \frac{3}{2} PV$  so the change in thermal energy due to the expansion is  $\Delta U = \frac{3}{2} P_f V_f - \frac{3}{2} P_i V_i = \frac{3}{2} (1.013 \times 10^5 \text{ Pa}) (3.1 \times 10^{-3} \text{ m}^3) - \frac{3}{2} (1.013 \times 10^5 \text{ Pa}) (1.0 \times 10^{-3} \text{ m}^3) = 1215.6 \text{ J}$  (11) The heat added to the gas must be enough to do the work plus increase the thermal energy, so  $Q = \Delta U + W = 1620.8 \text{ J}$  (12) Example 2.

## **COMPRESSION WORK An**



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Thermodynamic identity (Schroeder 3.4)  
 $dU = TdS - pdV$  (1.5) Total differentials When we have a total differential, the things in front of the  $dS, dV$ , etc are partial derivatives.  
 $T = \frac{\partial U}{\partial S} \quad -p = \frac{\partial U}{\partial V}$  (1.6) Also, you can integrate along a path using a total differential, and can do linear algebra using total differential equations, e.g. substituting one for another.

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